

# **WATER 390/590: Water Chemistry and Analysis**

**Fall Semester 2016**

## **SYLLABUS**

### **Course Information:**

Lecture Time: Monday/Wednesday/Friday 11:00 am – 11:50 am

Lecture Location: 255 Trainer Natural Resources Building

Credits: 4

Lab Times:

Section 1 – Thursday 9:00 am – 10:50 am

Section 2 – Thursday 1:00 pm – 2:50 pm

Section 3 – Tuesday 2:00 pm – 3:50 pm

Lab Location: 261 Trainer Natural Resources Building

Prerequisite: WATR 389, CHEM 106 or 117, and CNR or Biology major

### **Instructor Information:**

Dr. Kyle Herrman

Email: Kyle.Herrman@uwsp.edu (*preferred contact method*)

Office: 263 Trainer Natural Resources Building

Office Phone: 715-346-4832

### **Office Hours:**

Time: Tuesday 12:00 pm - 2:00 pm

Location: 263 Trainer Natural Resource Building

Or by appointment if the assigned hours do not work

### **Course Objective:**

The objective of this class is to expose students to the principles of water chemistry in human dominated landscapes. This will be accomplished using direct instruction methods during lecture and hands-on experience in the lab and in the field. After completing this course a student will be able to interpret the water chemistry data from an aquatic ecosystem and be able to properly collect, prepare, and process water samples for analysis. We will cover a variety of topics ranging from thermodynamics to unit conversion to carbonate chemistry so it is vital that students stay up to date on lecture topics and seek help if they are unsure of any course material. DO NOT wait until the last minute to get help because all of the material we will cover throughout the semester is comprehensive.

Learning objectives:

- Demonstrate the importance of abiotic and biotic impacts on water chemistry
- Convert common units used in water chemistry evaluation
- Properly collect, process, preserve, and analyze water samples
- Recognize the role of water chemistry and how it is used to evaluate aquatic ecosystems

### Required text:

None. The book assigned at the bookstore is a text that will help you with basic chemistry concepts if you need a refresher.

### Grades:

Scale:

A	93-100	C	73-76
A-	90-92	C-	70-72
B+	87-89	D+	67-69
B	83-86	D	63-66
B-	80-82	D-	60-62
C+	77-79	F	<60

Assignments:

	<u>Points</u>	<u>Total</u>	<u>Percent of Total Grade</u>
Participation	15	15	7.5%
Homework (4)	5	20	10%
Exams (2)	35	70	35%
Lab Report			
Outline	5	5	2.5%
Introduction	5	5	2.5%
Peer review	5	5	2.5%
Methods & Stats	10	10	5%
Results & Figs/Tables	10	10	5%
Full paper	60	60	30%

### Participation:

These points will be assessed on attendance in lecture and lab and on participation during in-class discussions. However, the bulk of these points will be assessed on your active participation during labs and the successful completion of assignments during labs.

### **Classroom Civility:**

Any successful learning experience requires mutual respect on the part of the student and the instructor. Neither instructor nor student should be subject to others' behavior that is rude, disruptive, intimidating, or demeaning. The instructor has primary responsibility for and control over classroom behavior and maintenance of academic integrity.

### **Homework:**

Assignments will be handed out throughout the semester that will address recently covered material. It may require calculations, diagrams, drawings, or using a spreadsheet to create a graph. Points will be assessed on three main criteria: 1) does the student answer the question(s); 2) is the student's argument clearly laid out and organized; 3) is the assignment legible and easy to read. Students can discuss concepts with each other but DO NOT turn in identical assignments. This includes graphs that are the same or verbiage that is identical to another student's assignment. If identical material is turned in no credit will be given to anyone involved.

### **Exams:**

Two take-home exams will be given in class and consist of essay/calculation questions. Exams will not be graded on a curve but partial credit will be given as long as the student clearly answers questions in an organized manner that I can follow. Direct comparison and working on specific calculations with other students is NOT allowed. You are free to discuss general approaches to problems with other students but YOU and YOU ALONE must solve each problem. If I notice that solutions from two or more students are too similar or if graphs look alike then I will take the appropriate steps to make sure all involved parties will not receive credit.

### **Lab Report:**

Based on the data collected throughout the semester from the three sites you will be expected to write a 6-page paper (including figures and tables) discussing water chemistry. This paper will have a typical scientific paper format as follows: Introduction, Methods, Results, and Discussion. It will discuss how the data from each site compares and how the data compares to other sites in the scientific literature. Sections will be turned in throughout the semester to receive feedback from the instructor and from peers. A minimum of 4 references from peer reviewed journal articles is required. For graduate students, the requirements increase to a 12 page paper and a minimum of 8 references. More details will be given later in the semester regarding format and style.

**Academic Misconduct:**

Violations of academic integrity will result in automatic failure of the class and referral to the proper university officials. Lab reports will be submitted on 2DL and will be analyzed for plagiarism via the program Turnitin. The work a student submits in class is expected to be the student's own work and must be work completed for that particular class and assignment. Students wishing to build on an old project or work on a similar topic in two classes must discuss this with the professor. Academic dishonesty includes but is not limited to: cheating on an examination and submitting an assignment as your own work when all or part of the assignment is the work of another without proper citation. Sanctions can be applied whether the violation was intentional or not so please know how to properly cite references for a scientific paper.

For further information regarding UWSP policy please refer to Chapter 14 in the University Handbook (<http://www.uwsp.edu/admin/stuaffairs/rights/rightsChap14.pdf>)

**Late Policy:**

Assignments are considered late if they are not turned in at the beginning of lecture on the due date. Assignments can be turned in late but 1 point will be taken off for each day the assignment is late. Exams must be turned in at the beginning of class on the day specified and will be deducted 1 letter grade per day until they are turned in.

**Attendance:**

If you are going to miss a lecture or an exam please contact me as soon as possible. If you have a documented absence then due dates can be extended. However, if you do not have an approved excuse for your absence then the appropriate late policies will be applied.

## Tentative Schedule (could change as semester progresses):

### Lecture Schedule

	Date	Lecture Topic
1	Sept 7	Syllabus and Water Basics
2	Sept 9	Common units and conversions
	Sept 12	
3	Sept 14	Dissolved Oxygen
	Sept 16	
4	Sept 19	Redox Reactions
	Sept 21	
	Sept 23	
	Sept 26	
5	Sept 28	How to write a scientific paper
	Sept 30	
6	Oct 3	Carbon Cycle
7	Oct 5	Nitrogen Cycle
	Oct 7	
8	Oct 10	Phosphorus Cycle
9	Oct 12	Nutrient Limitations
10	Oct 14	Mass balances in aquatic ecosystems
	Oct 17	
	Oct 19	Review ( <i>Exam I assigned</i> )
	Oct 21	No Class
	Oct 24	Poisoned Waters Part I
	Oct 26	Poisoned Waters Part II ( <i>Exam I due</i> )
11	Oct 28	Thermodynamics
	Oct 31	
12	Nov 2	Acid/Base Chemistry
	Nov 4	
	Nov 7	
	Nov 9	
	Nov 11	
13	Nov 14	Carbonate Chemistry
	Nov 16	
	Nov 18	
	Nov 21	
	Nov 23	NO CLASS
	Nov 25	NO CLASS
14	Nov 28	Precipitation/Dissolution
	Nov 30	
	Dec 2	
15	Dec 5	Organic Pollutants
	Dec 7	
	Dec 9	
	Dec 12	Review ( <i>Exam II assigned</i> )
	Dec 14	No Class ( <i>Paper due Dec 15 by 5:00 pm</i> )
<i>Finals Week</i>		
Dec 19 from 10:15 am – 12:15 pm: <i>Exam II due</i>		

**Lab Schedule (will most likely change based on field conditions)**

	<b>Week of:</b>	<b>Lab Topic</b>
	Sept 5	NO LAB
1	Sept 12	Calibrating Hydrolab's
2	Sept 19	Collect Field Samples and <i>In situ</i> Data
3	Sept 26	Filter samples/TSS/Alkalinity
4	Oct 3	Using Excel
5	Oct 10	Standards and calibration curves
6	Oct 17	Discuss outline
7	Oct 24	FIA – SRP and persulfate digestion for TN and TP
8	Oct 31	Data analysis/Statistics
9	Nov 7	Ion Chromatograph – Anions (Cations will be run later)
10	Nov 14	DOC/DIC
	Nov 21	NO LAB
11	Nov 28	Atomic absorption – Metals analysis
12	Dec 5	Log C-pH Diagrams
	Dec 12	Workshop for paper